PSBB LEARNING LEADERSHIP ACADEMY

SCIENCE STANDARD VIII

MID TERM REVISION 2 WORKSHEET

1. WRITE THE BALANCED CHEMICAL EQUATION:

a. Reaction between calcium and sulphuric acid.
Ans. $Ca + H_2SO_4 = CaSO_4\ H_2$
b. Reaction between iron and copper sulphate.
Ans. $Fe + CuSO_4 = FeSO_4 + Cu$
c. Reaction between potassium and water.
Ans. $K + 2H_2O = 2KOH + H_2$

2. COMPLETE THE FOLLOWING REACTION AS BALANCED CHEMICAL EQUATION:

a. $Ca + 2HCl = CaCl_2 + H_2$
b. $Na + 2H_2O = 2NaOH + H_2$
c. $H_2 + AgNO_3 = H_2NO_3 + Ag$
d. $Zn + CuCl_2 = ZnCl_2 + Cu$

3. DEFINE THE FOLLOWING:

a. Galvanisation
Ans. Process of coating metals with zinc to protect it from rusting.
b. Rusting
Ans. When iron reacts with oxygen in presence of moisture, it forms a reddish brown power, Iron Oxide, Fe_sO_3.

4. ANSWER THE FOLLOWING:

1. With the help of two examples, explain how electroplating process is useful to us.
Ans. Electroplating helps us in many ways to prevent metals from rusting, For eg. Electroplating is used in gates which are exposed to oxygen in presence of moisture but it does not rust because its electroplated, Same with car wheel rims.

2. What will happen when
    a. Calcium is reacting with dil. HCl.

Ans. Calcium displaces Hydrogen in Hydrochloric acid because calcium is more reactive than hydrogen. So it forms Calcium Chloride. $2Ca + 2HCL = 2CaCl + H_2$ (**BALANCED EQUATION**)

b. Magnesium burns in air

Ans. Magnesium burns in air to form magnesium oxide, $2Mg + O_2 = 2MgO$ (**BALANCED EQUATION**)